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following equation in standard form $0.5x + 4.8y = 10$ this equation is not in standard form because the coefficients are not integers to make them integers we can multiply each term

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elements exhibit only one normal oxidation state

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web a tank is charged with a mixture of 1.0×10^3 mol of oxygen and 4.5×10^3 mol of helium calculate the mole fraction of each gas in the mixture solution the given parameters are $n_{\text{he}} = 4.5 \times 10^3$ mol and $n_{\text{o}_2} = 1.0 \times 10^3$ mol mole fraction can be calculated as $x_{\text{he}} = \frac{4.5 \times 10^3 \text{ mol}}{4.5 \times 10^3 \text{ mol} + 1.0 \times 10^3 \text{ mol}} = \frac{4.5}{5.5}$

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